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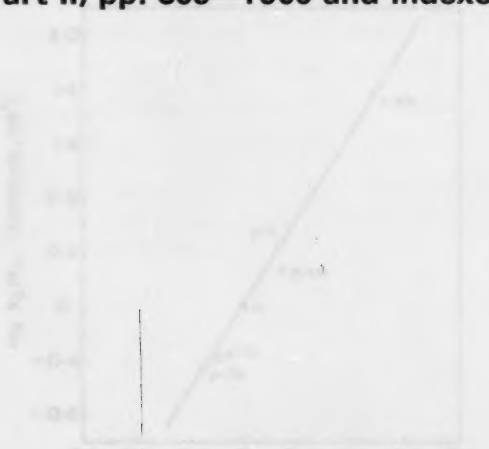

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ERRATA. Section B

1961, page 4923. The values of k and k_{rel} in the Table *refer* to a temperature of 50·0°, *not* 25° as stated.

1965, page 4667, line 27. *For* $[\alpha]_D$ −19·1 *read* $[\alpha]_D$ +19·1.

1968, page 310, left hand column, line 22. *Replace* ' In no case was coupling between ring and side chain observed in the ^{19}F spectrum; ' *by* ' With the exceptions of perfluorotoluene, methylpentafluorobenzene and ethylpentafluorobenzene, no coupling was observed in the ^{19}F spectrum between ring and side-chain; '.

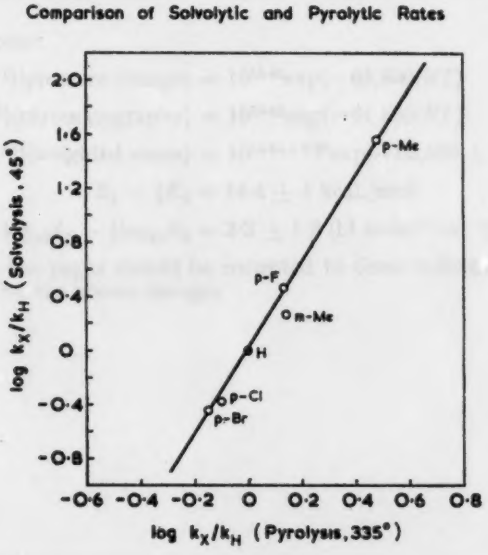
1968, page 527, right hand column, lines 2 and 6. *For* α *read* β. Lines 3 and 4. *For* β *read* α.

1968, page 633, right hand column, Formulae. *For* $\overset{+}{H}OKH\cdot NMe_2$ *read* $\overset{+}{H}OCN(Me)_2$.

1968, pages 760—762. In the Title, Summary and text *for* 1,4,7,96-tetra-azaphenaline *read* 1,4,7,9b-tetra-azaphenaline.

1968, page 769, right hand column, Formulae. *Interchange* formulae A and B.

1968, page 808, left hand column. *Insert* Figure 3.



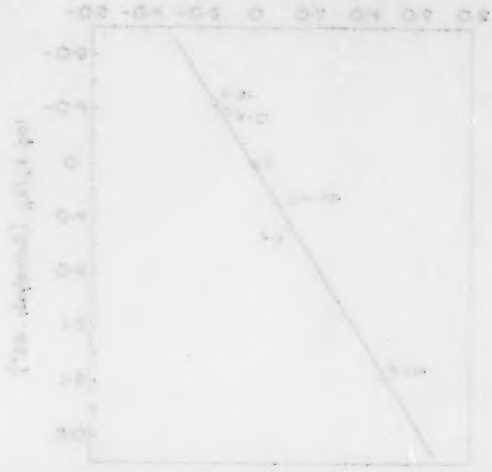
1968, page 844, Table 10, heading of column. *For* $10^3 l_{LA}$ *read* $10^3 k_{LA}$.

page 848, Equation (13) should *read*

$$\dddot{Y} = (\dddot{y}_A \dddot{y}_{HA}/\dddot{y}_{+(H)})_n = \text{etc.}$$

1968, page 851, right hand column, line 16. *For* perchloric acid *read* hydrochloric acid.

page 855, left hand column. The last factor in the equation following equation (12) should *read* $\dfrac{\ddot{Y}_D}{\ddot{Y}_H \ddot{Y}_D}$



Errata

1968, page 1036, The heading to Table 1 should *read* – Effects of substituents on rates of cleavage of $X{\cdot}C_6H_4{\cdot}SnMe_3$ compounds at 50·0° –.

1968, page 1280, Table 1. Delete the following – at 40 ± 0·01°.

Correction to Paper Entitled 'A Kinetic Study of the Gas-phase Thermal Decomposition of 1,1-Dimethyl-1-silacyclobutane'

By **M. C. Flowers** and **L. E. Gusel'nikov**, *J. Chem. Soc.* (*B*), 1968, 419

All temperatures were incorrectly calculated. Table 2 should read:

Temp.	10^4k_1(sec.$^{-1}$) † Pressure data	Temp.	10^4k_1(sec.$^{-1}$) † Chromatography	$10^5k_2/k_2^{\frac{1}{2}}$ l.$^{\frac{1}{2}}$ mole$^{-\frac{1}{2}}$ sec.$^{-\frac{1}{2}}$)
450·1°	6·68(6)	450·7°	5·95(5)	10·25
443·5	4·58(2)	437·1	2·47(2)	8·01
436·6	3·00(2)	436·0	2·21 *	
430·0	1·96(2)	423·0	1·16(7)	6·88
423·7	2·325(2)	394·3	0·149(2)	
416·2	0·822(1)			
410·1	0·483(3)			
400·9	0·277(2)			
394·2	0·180(2)			

* Obtained in packed reaction vessel. † Number of determinations are given in parentheses.

The Arrhenius equations become:

$$k_1(\text{sec.}^{-1})(\text{pressure change}) = 10^{15\cdot81}\exp(-62{,}800/RT)$$

$$k_1(\text{sec.}^{-1})(\text{chromatography}) = 10^{15\cdot43}\exp(-61{,}800/RT)$$

$$k_1(\text{sec.}^{-1})(\text{weighted mean}) = 10^{15\cdot64\pm0\cdot30}\exp[-62{,}500 \pm 800/RT]$$

$$E_2 - \tfrac{1}{2}E_3 = 14{\cdot}4 \pm 4 \text{ kcal./mole}$$

$$\log_{10}A_2 - \tfrac{1}{2}\log_{10}A_3 = 3{\cdot}3 \pm 1{\cdot}2 \text{ (l.}^{\frac{1}{2}}\text{ mole}^{-\frac{1}{2}}\text{ sec.}^{-1}\text{)}$$

All temperatures referred to in the paper should be corrected to those indicated in the revised table. The conclusions reached are unaffected by the above changes.

[7/1487/A. September 1968]

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